Using the materials provided, first cut out the step number and procedure, then put the steps in order for the acidic/basic oxidation number method.

Add OH- (aq) to both sides equal in number to the number of H+(aq) present.

Using the chemical formulas, determine the number of electrons transferred per reactant. Use formula subscripts to do this.)

Calculate the simplest whole number of coefficients for the reactants that will balance the total number of electrons transferred. Balance the reactants and products.

Combine H+(aq) and OH-(aq) on the same side to formH2O(l) and cancel the same number of H2O(l) on both sides.

Using the change in oxidation numbers, write the number of electrons transferred per atom.

Balance the O atoms using H2O(l), and then balance the H atoms using H+(aq).

Assign oxidation numbers and identify the atoms/ions whose oxidation number change.

Step 1

Step 2

Step 3

Step 4

Step 5

Step 6

Step 7

For basic solutions only